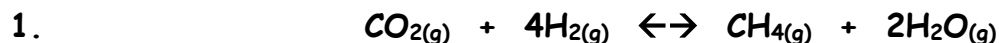


All ICE Charts

0.10 mol of CO_2 and 0.20 mol of H_2 are placed in a 1 L container. At equilibrium the concentration of H_2 is 0.10 mol/L. $K_{\text{eq}} = ?$



0.10 mol/L of HOCl is allowed to react. What is the concentration of Cl_2O at equilibrium?



0.0010 mol of bromine and 1.00 mol of fluorine are added to a 1 L container. What is the concentration of BrF_5 at equilibrium?



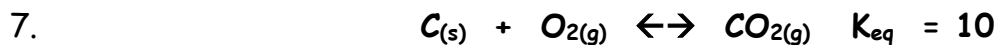
0.10 mol PCl_5 , 0.20 mol PCl_3 and 0.30 mol Cl_2 are added to a 1 L container. What is the concentration of Cl_2 at equilibrium?



The starting concentration of HCl is 1.0M. What is the equilibrium concentration of H_2 ?



An equilibrium mixture contains 0.10 M CO , 0.00050M Cl_2 and 0.10M CoCl_2 . What is the new equilibrium concentration of CO after 0.050 mol of COCl_2 are removed from 1 L of the mixture?



Into a 500 mL container, 0.50 mol of C and 0.50 mol of O_2 are added. What is the concentration of CO_2 at equilibrium?



0.10 M C is added to a container. What is the concentration of A at equilibrium?